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The name of a base is the name of the cation followed by the name of the anion (hydroxide). The formula of a base is written by showing the number of hydroxide ions needed to balance the positive charge on the cation. After reading Lesson 9.4, answer the following questions. Names and Formulas of Acids hydrogen1.

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86 Guided Reading and Study Workbook 6. Complete the table using the law of multiple proportions. Practicing Skills: Naming Chemical Compounds (pages 276–277) 7. How can a flowchart help you to name chemical compounds? 8. Use the flowchart in Figure 9.20 on page 277 to write the names of the following compounds: a. CsCl _____

Name Date Class CHEMICAL NAMES AND FORMULAS 9

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Chemical Formula Writing Worksheet Two Write chemical formulas for the compounds in each box. The names are found by finding the intersection between the cations and anions. Example: The first box is the intersection between the “zinc”

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cation and the “chloride” anion, so you should write “ZnCl₂”, as shown. Cations

STUDY GUIDE: Naming & Formulas of Ionic Compounds

The property that is useful for keeping track of heat transfers in chemical enthalpy and physical processes at constant pressure is called _____. Chapter 17 Thermochemistry 185

05_Chem_GRSW_Ch17.SE/TE 6/11/04 3:49 PM Page 186

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chemical energy Guided Reading and Study Workbook. Science Explorer Grade 6 GSW-6_01/FNL 5/2/01 12:27 PM Page 9 Name _____ Date _____ Class _____ Measuring Matter SECTION 1-2 (pages 34-39) This section explains the difference between mass and weight. It also explains what the density of a substance is. Mass

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94 Guided Reading and Study Workbook 1.00 mol 6.02 10²³ particles 6.02 10²³ particles 1.00 mol molar mass 1.00 mol 9. What is Avogadro’s hypothesis? 10. Look at Figure 10.9 on page 300 to help you answer this question.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

14 Study Guide for An Introduction to Chemistry Exercise 2.3 - Cations and Anions: Identify each of the following as a cation or an anion, and determine the charge on each. (Obj 22) a magnesium atom with 12 protons and 10 electrons (+12) + (10) = +2 This is a +2 cation. a fluorine atom with 9 protons and 10 electrons

Chapter 2 The Structure of Matter and the Chemical Elements

Chemical Equations. A chemical equation is a written expression of a chemical reaction; e.g., $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ Reactants are written on the left, and products are written on the right. In a balanced equation the total numbers of atoms of each kind on both sides are the same.

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3A: Compounds, Naming, Reaction Equations, and Formula ...

118 Study Guide for An Introduction to Chemistry Section 9.2 Relating Mass to Number of Particles Goals To show how to do a procedure called counting by weighing. To introduce atomic mass and show how it can be used to convert between the mass of a sample of an element and the number of atoms that the sample contains. Even a tiny sample of an element contains a huge number of atoms.

Chapter 9 Chemical Calculations and Chemical Formulas

Name_____ Class_____ Date_____ Chapter 2 The Chemistry of Life ... how atoms of the same element can have a different number of neutrons and describes the two main types of chemical bonds. Atoms (page 35) 1. The basic unit of matter is called a(an) . 2. Describe the nucleus of an atom. ... Guided Reading and Study Workbook/Chapter 2 21

The Chemistry of Life

Chemical Names and Formulas Planning Guide 9.1 Naming Ions, pp. 253-259 1 block or 2 periods 9.1.1 Identify the charges of monatomic ions by using the periodic table, and name the ions. 9.1.2 Define a polyatomic ion and write the names and formulas of the most common polyatomic ions. 9.1.3 Identify the two common endings for the names of

9 Chemical Names and Formulas Planning Guide Lesson Plan ...

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PRENTICE HALL SCIENCE EXPLORER Grade 8

The Chemical Names and Formulas chapter of this Prentice Hall Chemistry Companion Course helps students learn the essential lessons associated with chemical names and formulas. Each of

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these simple...

Prentice Hall Chemistry Chapter 9: Chemical Names and

...

Name each quantity that is represented in the equation for heat change in an aqueous solution. q ΔH m C ΔT Thermochemical Equations (pages 514-517) 7. What happens to the temperature of water after calcium oxide is added? 8. A chemical equation that includes the heat change is called a _____ equation. 9.

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