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10 Marking scheme: End-of-chapter test. 1 a Element Carbon Hydrogen Oxygen Percentage 62.1 10.3 27.6. Mass in 100 g 62.1 g 10.3 g 27.6 g. Number of moles. $12.0 \text{ g} / 62.1 \text{ g} = 5.18$ $1.0 \text{ g} / 10.3 \text{ g} = 10.3$ $16.0 \text{ g} / 27.6 \text{ g} = 1.73$ [1] Relative number of atoms 1.73 5.18 = 3 1.73 10.3 = 6 1.73 1.73 = 1.

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13 Marking scheme: End-of-chapter test. 1 a. $C_nH_{2n+1}OH$ [1] cButan-1-ol and 2-methylpropan-2-ol [1] structural isomers [1] dCompounds with the same molecular formula [1] but different structural/displayed formulae [1] e2-methylbutan-2-ol [1] 2 a $C_6H_{12}O_6 \rightarrow 2CH_3CH_2OH + 2CO_2$ [1] fermentation [1]

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strip A, i.e. the 1 a (The force F is given by $F = BIl \sin \theta$.) strips attract each other.

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